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Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point.

Express the heat in kilojoules to three significant figures.

$q =$

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Correct

At its boiling point, the enthalpy of vaporization (ΔH_{vap}) of water is 40.7 kJ/mol. Once the given mass, 2.58×10^3 g, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by ΔH_{vap} yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25 °C condenses on the surface of a 55-g block of aluminum that is initially at 25 °C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, C_{Al} , is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

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